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8.4 -5-17 -22 = $x \times 10 = 2.05 \times 10$. 17. 4.1 $\times 10$. 4.1. b) 5.4 $\times 10$. 7.3 $\times 10$. 32. 14 = 5.4. 7.3. $x \times 10$. 32-14 = .74 $\times 10$. 18 = 7.4 $\times 10$. 17. Information: Adding and Subtracting Using Scientific Notation. Whenever you add or subtract two numbers in scientific notation, you must make sure that they have the same exponents. Your answer will then have the ...

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Chem Quest - Mr. Smith

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3. Iodine has 17 more electrons than krypton and so you can begin by writing [Kr] followed by orbitals for 617 more electrons. After 4p comes 5s2... Iodine = [Kr]5s24d105p5 Any of the noble gases can be used for abbreviations. Here are a few more examples: iron = [Ar]4s23d6 cesium = [Xe]6s1 phosphorus = [Ne]3s23p3

ChemQuest 15

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Given your answer to number 7, you should be able to answer the following questions. a) Find the pH of a solution that has a pOH of 4.6. 9.4; pH + pOH = 14 pH = 14 - pOH = 9.4 b) The [H³O⁺] of a solution is 2.55 $\times 10^{-3}$ M.

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7; since chlorine is in column 17 (or "7A") it has 7 valence electrons. 3. Since CCl₄ is made up of one carbon and four chlorine atoms, how many total valence electrons does CCl₄ have? Hint: add your answer to question 1 and four times your answer to question 2. $4 + 4(7) = 4 + 28 = 32$ 4. Verify that there are 32 electrons pictured in the ...

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Download Ebook Chemquest 49 Answer Key Chemquest 49 Answer Key The Online Books Page: Maintained by the University of Pennsylvania, this page lists over one ... Iodine is 80% 1271 17% 1261, and 3%1281. Calculate the average atomic mass of iodine. 2. Calculate the average atomic mass of lithium, which occurs as two isotopes that

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24 ChemQuest 8 Name: _____ Date: _____ Hour: _____ Information: Structure of the Atom Note the following symbols: (they are not to scale) = proton (positive charge) = electron (negative charge) = neutron (no charge) The following three diagrams are hydrogen atoms: 1 1H 2 1H 3 1H The following three diagrams are carbon atoms: Notice the type of notation used for atoms: 12 13 14 6C 6 C 6C (6 ...

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Calculate the average atomic masses. Round all answers to two decimal places. Record your answer with the correct units. 1. Iodine is 80% 1271 17% 1261, and 3%1281. Calculate the average atomic mass of iodine. 2. Calculate the average atomic mass of lithium, which occurs as two isotopes that have the following

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118 ChemQuest 39 Name: _____ Date: _____ Hour: _____ Information: Molarity Concentration is a term that describes the amount of solute that is dissolved in a solution. Concentrated solutions contain a lot of dissolved solute, but dilute solutions contain only a little.

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